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Titles

Hawley's Condensed Chemical Dictionary (14th Edition)

Handbook of Refractory Carbides and Nitrides

Publisher: William Andrew Publishing/Noyes

Author/Editor: : Pierson, H.O.

ISBN: 0-8155-1392-5

Copyright: © 1996

Knovel Release: April 13, 2001

Description: Refractory carbides and nitrides are useful materials with numerous industrial applications and a promising future, in addition to being materials of great interest to the scientific community. Although most of their applications are recent, the refractory carbides and nitrides have been known for over one hundred years.

Ordering Info

The industrial importance of the refractory carbides and nitrides is growing rapidly, not only in the traditional and well-established applications based on the strength and refractory nature of these materials such as cutting tools and abrasives, but also in new and promising fields such as electronics and opto-electronics.

Handbook of Carbon, Graphite, Diamond and Fullerenes - Properties, Processing and Applications

Dictionary of Biochemistry and Molecular Biology (2nd Edition)

Harper's Illustrated Biochemistry (26th Edition)

Polymer Blends Handbook, Volumes 1-2

Encyclopedia of Molecular Biology, Volumes 1-4

Chemistry of the Elements (2nd Edition)

Sol-Gel Technology for Thin Films, Fibers, Preforms, Electronics and Specialty Shapes

Chemistry of Nanostructured Materials NEW

Ceramic Interfaces - Properties and Applications NEW

Springer Handbook of Nanotechnology

Encyclopedia of Controlled Drug Delivery, Volumes 1-2

Clinician's Handbook of Prescription Drugs

Hydrogen Degradation of Ferrous Alloys

material is sodium chloride (NaCl). Other examples of ionic bond are the salt-like carbides described in Sec. 5.4.

5.2 Covalent Bond

Covalent bonds are formed by the sharing of electrons (rather than transfer). Typically two atoms share a pair of electrons. A covalent structure, that of diamond, is shown schematically in Fig. 2.1.[7] The shaded regions designate a high probability of finding the shared electrons (see Sec. 3.3).

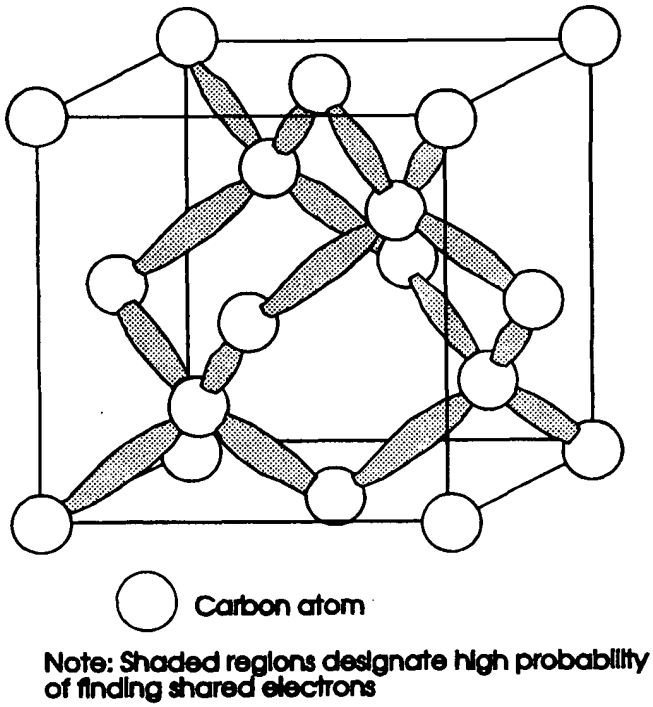

Figure 2.1: Schematic representation of the structure of the diamond crystal.